Chemistry Colligative Properties Of Solutions Section Review

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Chapter 13 Properties of Solutions Chemistry, The Central Science , 10th edition Theodore L. Brown; H. Eugene LeMay, Jr.; and Bruce E. Bursten. Solutions Solutions Colligative Properties • Changes in colligative properties depend only on the number of solute particles present, not on the identity of

CHAPTER 13 REVIEW Ions in Aqueous Solutions and Colligative Properties MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Match the four compounds on the right to their descriptions on the left.

Colligative Properties- Page 1 Lecture 4: Colligative Properties • By definition a colligative property is a solution property (a property of mixtures) for which it is the amount of solute dissolved in the solvent matters but the kind of solute does not matter.

CHAPTER 14 REVIEW Ions in Aqueous Solutions and Colligative Properties SECTION 14-2 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. a. Predict the boiling point of a 0.200 m solution of glucose in water. b.

WORKSHEET:SOLUTIONS AND COLLIGATIVE PROPERTIES SET A: 1. Find the molarity of all ions in a solution that contains 0.165 moles of aluminum chloride in 820. ml solution. Answer: [Al 3+]= 0.201 M , (Cl-] = 0.603M. 2. Find the molarity of each ion present after mixing 27 ml of 0.25 M HNO3 with 36 ml of 0.42 M Ca(NO3)2

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Chapter 16: Colligative Properties of Solutions 45 16-4. The mole fraction of (NH 4) 2SO 4(aq) is given by x (NH 4) 2SO 4 = n (NH 4) 2SO 4 n (NH 4) 2SO 4 +n H 2O Because (NH 4) 2SO 4(aq) is a strong electrolyte, it dissociates completely into NH + 4 (aq) and SO2– 4 (aq) ions. Assume a one kilogram solution. The number of moles of ions in one

Chapter 16 Solutions 403 Section Review Objectives • Identify the three colligative properties of solutions • Describe why the vapor pressure, freezing point, and boiling point of a solution differ from those properties of the pure solvent. Vocabulary • colligative properties • freezing-point depression • boiling-point elevation Part ACompletion Use this completion exercise to check

SECTION 16.3 COLLIGATIVE PROPERTIES OF SOLUTIONS (pages 487–490) This section explains why a solution has a lower vapor pressure, an elevated boiling point, and a depressed freezing point compared with the pure solvent of that solution. Vapor Pressure Lowering (pages 487–488) 1. Properties of a solution that depend only on the number of

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